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Homogeneous chemical reaction on unsteady flow past an accelerated vertical plate with variable temperature and thermal radiation

R. Muthucumaraswamy¹, P. Balachandran², V. Prasanth³ and K. Ganesan⁴

¹Department of Applied Mathematics, Sri Venkateswara College of Engineering, Sriperumbudur 602 105, India,

²Department of Mathematics, SRM Valliammai Engineering College, Potheri, Kattankulathur 603203.

³Department of Chemical Engineering, Sri Venkateswara College of Engineering, Sriperumbudur 602 105, India

⁴Department of Mathematics, SRM University, Potheri, Kattankulathur 603203.

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ABSTRACT

An exact analysis of unsteady radiative flow past a uniformly accelerated infinite vertical plate with variable temperature with uniform mass diffusion is presented here, taking in to account the homogeneous chemical reaction of first order. The plate temperature is raised

linearly with time and the concentration level near the plate is also raised to C'_w . The dimensionless governing equations are solved using Laplace-transform technique. The velocity, temperature and concentration fields are studied for different physical parameters like thermal Grashof number, mass Grashof number, Schmidt number, Prandtl number, radiation parameter, chemical reaction parameter and time. It is observed that the velocity increases with increasing values of thermal Grashof number or mass Grashof number. But the trend is just reversed with respect to the thermal radiation parameter. It is also observed that the velocity increases with decreasing chemical reaction parameter.

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Introduction

Thermal radiation effects with heat and mass transfer play an important role in manufacturing industries for the design of fins, steel rolling, nuclear power plants, gas turbines and various propulsion device for aircraft, combustion and furnace design, materials processing, energy utilization, temperature measurements, remote sensing for astronomy and space exploration. England and Emery [6] have studied the thermal radiation effects of an optically thin gray gas bounded by a stationary vertical plate. Radiation effect on mixed convection along an isothermal vertical plate was studied by Hossain and Takhar[9]. Raptis and Perdikis [11] studied the effects of thermal radiation and free convection flow past a moving vertical plate. The governing equations were solved analytically. Das *et al*[5] have analyzed radiation effects on flow past an impulsively started infinite isothermal vertical plate. The dimensionless governing equations were solved by the usual Laplace-transform technique.

Chemical reactions can be classified as either heterogeneous or homogeneous processes. This depends on whether they occur at an interface or as a single phase volume reaction. In well-mixed systems, the reaction is heterogeneous, if it takes place at an interface and homogeneous, if it takes place in solution. In most cases of chemical reactions, the reaction rate depends on the concentration of the species itself. A reaction is said to be of first order, if the rate of reaction is directly proportional to the concentration itself. Chambre and Young [2] have analyzed a first order chemical reaction in the neighborhood of a horizontal plate. Das *et al*[3] have studied the effect of homogeneous first order chemical reaction on the flow past an impulsively started infinite vertical plate with uniform heat flux and mass transfer. Again, mass transfer effects on moving isothermal vertical plate in the presence of chemical reaction studied by Das *et al*[4]. The

dimensionless governing equations were solved by the usual Laplace-transform technique.

Gupta *et al* [7] studied free convection on flow past a linearly accelerated vertical plate in the presence of viscous dissipative heat using perturbation method. Kafousias and Raptis[10] extended the above problem to include mass transfer effects subjected to variable suction or injection. Free convection effects on flow past an exponentially accelerated vertical plate was studied by Singh and Naveen Kumar[12]. The skin friction for accelerated vertical plate has been studied analytically by Hossain and Shayo [8]. Mass transfer effects on flow past a uniformly accelerated vertical plate was studied by Soundalgekar[14]. Again, mass transfer effects on flow past an accelerated vertical plate with uniform heat flux was analyzed by Singh and Singh [13]. Basant Kumar Jha and Ravindra Prasad [1] analyzed mass transfer effects on the flow past an accelerated infinite vertical plate with heat sources.

Hence, it is proposed to study the first order chemical reaction on unsteady flow past a uniformly accelerated infinite vertical plate with variable temperature and uniform mass diffusion, in the presence of thermal radiation. The dimensionless governing equations are solved using the Laplace-transform technique. The solutions are in terms of exponential and complementary error function. Such a study found useful in chemical process industries such as wire drawing, fiber drawing, and food processing and polymer production.

Mathematical formulation

Here the unsteady radiative flow of a viscous incompressible fluid past a uniformly accelerated isothermal infinite vertical plate with variable mass diffusion in the presence of chemical reaction of first order has been considered. Here the unsteady flow of a viscous incompressible fluid which is initially at rest and surrounds an infinite vertical plate with

temperature T_∞ and concentration C'_∞ . The x -axis is taken along the plate in the vertically upward direction and the y -axis is taken normal to the plate. At time $t' \leq 0$, the plate and fluid are at the same temperature T_∞ . At time $t' > 0$, the plate is accelerated with a velocity $u = \frac{u_0^3 t'}{\nu}$ in its own plane and the temperature from the plate is raised linearly with time and the concentration level near the plate is also raised to C'_w . It is assumed that the effect of viscous dissipation is negligible in the energy equation and there is a first order chemical reaction between the diffusing species and the fluid. The fluid considered here is a gray, absorbing-emitting radiation but a non-scattering medium. Then under usual Boussinesq's approximation the unsteady flow is governed by the following equations:

$$\frac{\partial u}{\partial t'} = g\beta(T - T_\infty) + g\beta^*(C' - C'_\infty) + \nu \frac{\partial^2 u}{\partial y^2} \quad (1)$$

$$\rho C_p \frac{\partial T}{\partial t'} = k \frac{\partial^2 T}{\partial y^2} - \frac{\partial q_r}{\partial y} \quad (2)$$

$$\frac{\partial C'}{\partial t'} = D \frac{\partial^2 C'}{\partial y^2} - K_l(C' - C'_\infty) \quad (3)$$

With the following initial and boundary conditions:

$$u = 0, \quad T = T_\infty, \quad C' = C'_\infty \quad \text{For all } y, t' \leq 0$$

$$t' > 0: \quad u = \frac{u_0^3 t'}{\nu}, \quad T = T_\infty + (T_w - T_\infty)At', \quad C' = C'_w \quad \text{at } y = 0 \quad (4)$$

$$u \rightarrow 0, \quad T \rightarrow T_\infty, \quad C' \rightarrow C'_\infty \quad \text{as } y \rightarrow \infty$$

The local radiant for the case of an optically thin gray gas is expressed by

$$\frac{\partial q_r}{\partial y} = -4a^* \sigma (T_\infty^4 - T^4) \quad (5)$$

It is assumed that the temperature differences within the flow are sufficiently small such that T^4 may be expressed as a linear function of the temperature. This is accomplished by expanding T^4 in a Taylor series about T_∞ and neglecting higher-order terms, thus

$$T^4 \cong 4T_\infty^3 T - 3T_\infty^4 \quad (6)$$

By using equations (5) and (6), equation (2) reduces to

$$\rho C_p \frac{\partial T}{\partial t'} = k \frac{\partial^2 T}{\partial y^2} + 16a^* \sigma T_\infty^3 (T_\infty - T) \quad (7)$$

On introducing the following non-dimensional quantities:

$$U = \frac{u}{u_0}, \quad t = \frac{t' u_0^2}{\nu}, \quad Y = \frac{y u_0}{\nu}, \quad \theta = \frac{T - T_\infty}{T_w - T_\infty},$$

$$Gr = \frac{g\beta\nu(T_w - T_\infty)}{u_0^3}, \quad C = \frac{C' - C'_\infty}{C'_w - C'_\infty}, \quad Gc = \frac{\nu g\beta^*(C'_w - C'_\infty)}{u_0^3},$$

$$R = \frac{16a^* \nu^2 \sigma T_\infty^3}{k u_0^2}, \quad Pr = \frac{\mu C_p}{k}, \quad Sc = \frac{\nu}{D}, \quad K = \frac{\nu K_l}{u_0^2} \quad (8)$$

in equations (1), (3) and (7) reduces to

$$\frac{\partial U}{\partial t} = Gr \theta + Gc C + \frac{\partial^2 U}{\partial Y^2} \quad (9)$$

$$\frac{\partial \theta}{\partial t} = \frac{1}{Pr} \frac{\partial^2 \theta}{\partial Y^2} - \frac{R}{Pr} \theta \quad (10)$$

$$\frac{\partial C}{\partial t} = \frac{1}{Sc} \frac{\partial^2 C}{\partial Y^2} - K C \quad (11)$$

The negative sign of K in the last term of the equation (11) indicates that the chemical reaction takes place from higher level of concentration to lower level of concentration.

The initial and boundary conditions in non-dimensional quantities are

$$U = 0, \quad \theta = 0, \quad C = 0 \quad \text{for all } Y, t \leq 0$$

$$t > 0: \quad U = t, \quad \theta = t, \quad C = 1 \quad \text{at } Y = 0 \quad (12)$$

$$U \rightarrow 0, \quad \theta \rightarrow 0, \quad C \rightarrow 0 \quad \text{as } Y \rightarrow \infty$$

The dimensionless governing equations (9) to (11), subject to the initial and boundary conditions (12), are solved by the usual Laplace-transform technique and the solutions are derived as follows:

$$\theta = \frac{t}{2} \left[\exp(2\eta\sqrt{Rt}) \operatorname{erfc}(\eta\sqrt{Pr} + \sqrt{at}) + \exp(-2\eta\sqrt{Rt}) \operatorname{erfc}(\eta\sqrt{Pr} - \sqrt{at}) \right]$$

$$- \frac{\eta Pr \sqrt{t}}{2\sqrt{R}} \left[\exp(-2\eta\sqrt{Rt}) \operatorname{erfc}(\eta\sqrt{Pr} - \sqrt{at}) - \exp(2\eta\sqrt{Rt}) \operatorname{erfc}(\eta\sqrt{Pr} + \sqrt{at}) \right] \quad (13)$$

$$C = \frac{1}{2} \left[\exp(2\eta\sqrt{KtSc}) \operatorname{erfc}(\eta\sqrt{Sc} + \sqrt{Kt}) + \exp(-2\eta\sqrt{KtSc}) \operatorname{erfc}(\eta\sqrt{Sc} - \sqrt{Kt}) \right] \quad (14)$$

$$U = t(1 + 2bd) \left[(1 + 2\eta^2) \operatorname{erfc}(\eta) - \frac{2\eta}{\sqrt{\pi}} \exp(-\eta^2) \right] + 2(d + e) \operatorname{erfc}(\eta)$$

$$- d \exp(bt) \left[\exp(2\eta\sqrt{bt}) \operatorname{erfc}(\eta + \sqrt{bt}) + \exp(-2\eta\sqrt{bt}) \operatorname{erfc}(\eta - \sqrt{bt}) \right]$$

$$- e \exp(ct) \left[\exp(2\eta\sqrt{ct}) \operatorname{erfc}(\eta + \sqrt{ct}) + \exp(-2\eta\sqrt{ct}) \operatorname{erfc}(\eta - \sqrt{ct}) \right]$$

$$- d(1 + bt) \left[\exp(2\eta\sqrt{Rt}) \operatorname{erfc}(\eta\sqrt{Pr} + \sqrt{at}) + \exp(-2\eta\sqrt{Rt}) \operatorname{erfc}(\eta\sqrt{Pr} - \sqrt{at}) \right]$$

$$+ \frac{d Pr Gr \eta \sqrt{t}}{\sqrt{R}} \left[\exp(-2\eta\sqrt{Rt}) \operatorname{erfc}(\eta\sqrt{Pr} - \sqrt{at}) - \exp(2\eta\sqrt{Rt}) \operatorname{erfc}(\eta\sqrt{Pr} + \sqrt{at}) \right]$$

$$+ d \exp(bt) \left[\exp(-2\eta\sqrt{Pr(a+b)t}) \operatorname{erfc}(\eta\sqrt{Pr} - \sqrt{(a+b)t}) \right]$$

$$+ \exp(2\eta\sqrt{Pr(a+b)t}) \operatorname{erfc}(\eta\sqrt{Pr} + \sqrt{(a+b)t}) \Big]$$

$$- e \left[\exp(-2\eta\sqrt{KtSc}) \operatorname{erfc}(\eta\sqrt{Sc} - \sqrt{Kt}) + \exp(2\eta\sqrt{KtSc}) \operatorname{erfc}(\eta\sqrt{Sc} + \sqrt{Kt}) \right]$$

$$+ e \exp(ct) \left[\exp(-2\eta\sqrt{Sc(K+c)t}) \operatorname{erfc}(\eta\sqrt{Sc} - \sqrt{(K+c)t}) \right]$$

$$+ \exp(2\eta\sqrt{Sc(K+c)t}) \operatorname{erfc}(\eta\sqrt{Sc} + \sqrt{(K+c)t}) \Big] \quad (15)$$

Where,

$$a = \frac{R}{Pr}, \quad b = \frac{R}{1 - Pr}, \quad c = \frac{KSc}{1 - Sc}, \quad d = \frac{Gr}{2b^2(1 - Pr)}, \quad e = \frac{Gc}{2c(1 - Sc)}$$

and $\eta = \frac{Y}{2\sqrt{t}}$.

Results and discussion

For physical interpretation of the problem, the numerical computations are carried out for different physical parameters Gr, Gc, Sc, R, Pr, K and t upon the nature of the flow and transport. The value of the Schmidt number Sc is taken to be 0.6 which corresponds to water-vapor. Also, the value of Prandtl

number Pr are chosen such that they represent air ($Pr = 0.71$). The numerical values of the velocity, temperature and concentration are computed for different physical parameters like Prandtl number, thermal Grashof number, mass Grashof number, chemical reaction parameter, radiation parameter, Schmidt number and time are studied graphically.

The effect of the velocity for different values of the radiation parameter ($R = 0.2, 8, 20$), $Gr=10$, $Gc=5$, $K=5$ and $t=0.4$ are shown in figure 1. The trend shows that the velocity increases with decreasing radiation parameter. It is observed that there is a down trend in velocity in the presence of high thermal radiation. Figure 2 illustrates the effect of velocity for different values of the chemical reaction parameter ($K = 0.2, 5, 10$), $Gr = Gc = 5$, $R = 10$ and $t = 0.2$. It is observed that the increase in the chemical reaction parameter leads to a fall in the velocity. This shows that the velocity decreases tremendously in the presence of chemical reaction.

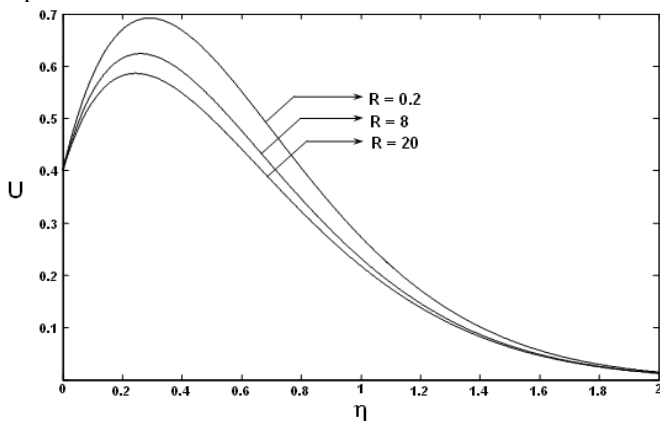


Fig. 1 Velocity profiles for different values of R

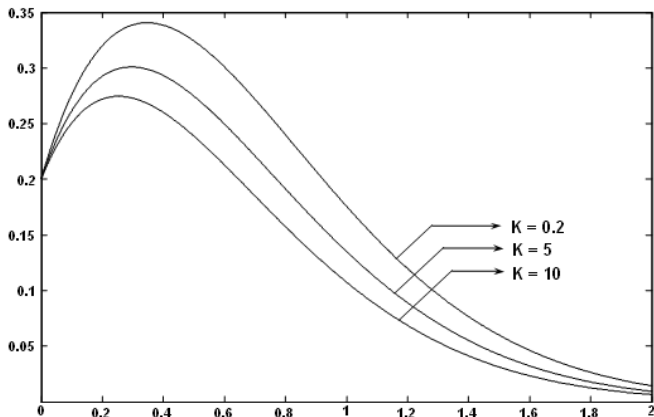


Fig. 2 Velocity profiles for different values of K

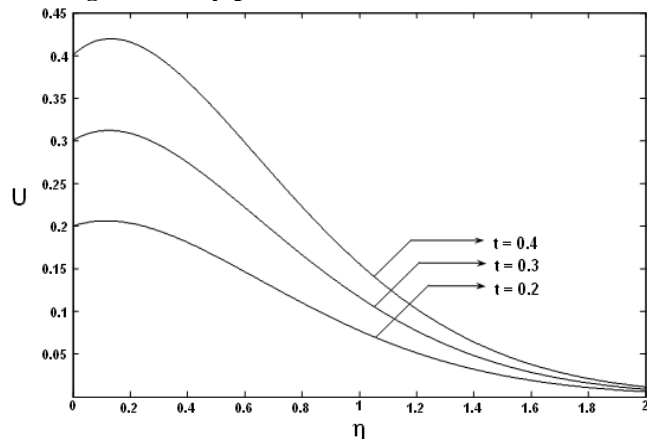


Fig. 3 Velocity profiles for different values of t

The velocity profiles for different values of ($t = 0.2, 0.3, 0.4$), $K = 0.2$, $R = 5$ and $Gr = Gc = 2$ are studied and presented in figure 3. It is observed that the velocity increases with increasing values of the time t . Figure 4 demonstrates the effect velocity fields for different thermal Grashof number ($Gr = 2, 10$), mass Grashof number ($Gc = 2, 5$), $K = 5$, $R = 10$ and $t = 0.4$. It is clear that the velocity increases with increasing values of the thermal Grashof number or mass Grashof number. It is quite interesting to note that the velocity increases very rapidly in the presence of mass Grashof number than thermal Grashof number.

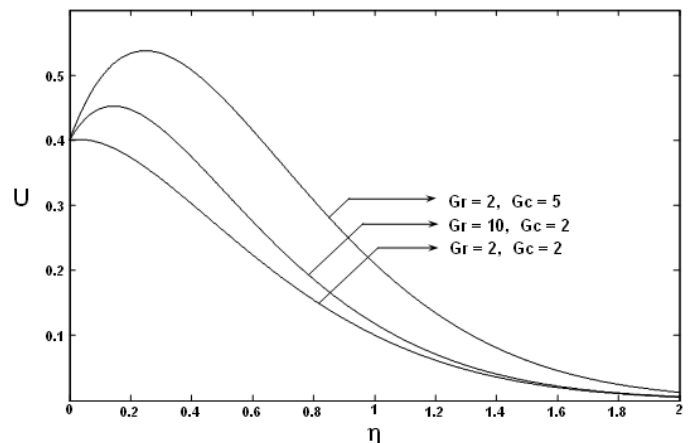


Fig. 4 Velocity profiles for different values of Gr and Gc

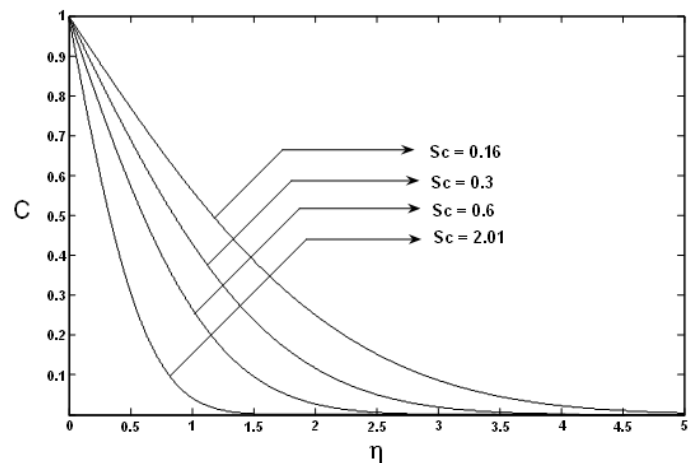


Fig. 5 Concentration profiles for different values of Sc

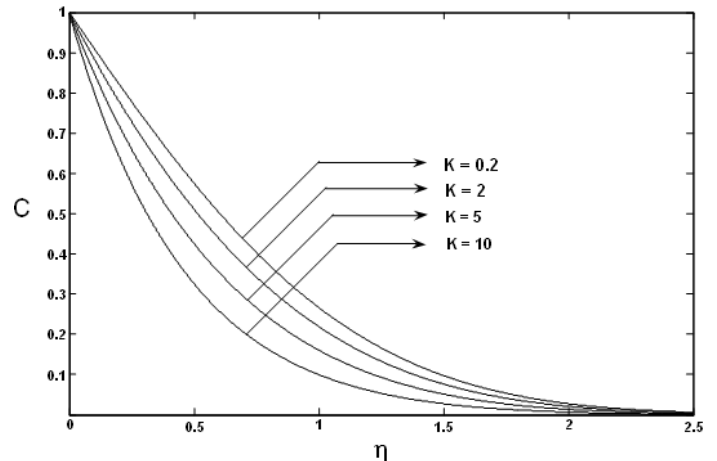


Fig. 6 Concentration profiles for different values of K

The effect of concentration profiles for different values of the chemical reaction parameter ($K = 0.2, 2, 5, 10$) and time $t = 0$ are presented in figure 5 in the presence of water vapour. The effect

of the chemical reaction parameter is dominant in concentration field. It is observed that the wall concentration increases with decreasing values of the chemical reaction parameter. Figure 6 represents the temperature profiles for different values of ($t=0.2, 0.4, 0.6, 1$) and $R=5$. The trend shows that the plate temperature increases with increasing values of the time t . The profiles have the common feature that the concentration decreases in a monotone fashion from the surface to a zero value far away in the free stream

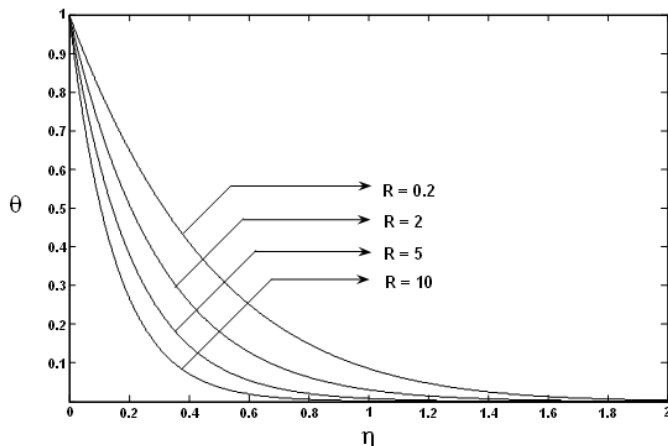


Fig. 7 Temperature Profiles for different values of R

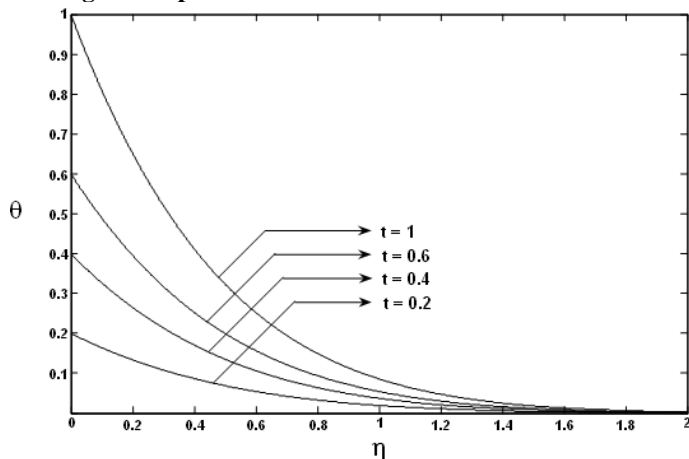


Fig. 8 Temperature Profiles for different values of t

The temperature profiles are calculated for different values of thermal radiation parameter ($R=2, 5, 7, 10$) and time $t=0.4$ from Equation (13) and these are shown in figure 7. The effect of thermal radiation parameter is important in temperature profiles. It is observed that the temperature increases with decreasing radiation parameter. Figure 8 illustrates the effect of temperature field for different values of time ($t=0.2, 0.4, 0.6, 1$) and $R=0.2$. It is observed that the plate temperature increase with increasing values of time t .

Conclusion

An exact solution of radiative flow past a uniformly accelerated infinite vertical plate with variable temperature and uniform mass diffusion, in the presence of homogeneous chemical reaction of first order has been studied. The dimensionless governing equations were solved by the usual

Laplace-transform technique. The effect of different physical parameters like chemical reaction parameter, radiation parameter, thermal Grashof number, mass Grashof number and t are studied graphically. It is observed that the velocity increases with increasing values of Gr, Gc and t . But the trend is just reversed with respect to the chemical reaction parameter or radiation parameter.

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