

## Synthesis of Cinnemaldine-aniline and Kinetic study

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### ABSTRACT

The second order reaction rate constant for the reaction of Cinnemaldehyde with aniline have been reported in ethanol in the temperature range 303 to 318 k. The rate of reaction is first order with respect to Cinnemaldehyde and first order with respect to aniline. Increase in temperature increases the rate of reaction. The thermodynamic parameters are used to explain the nature of reaction. Suitable reaction mechanism has been suggested for the formation of the Schiff base.

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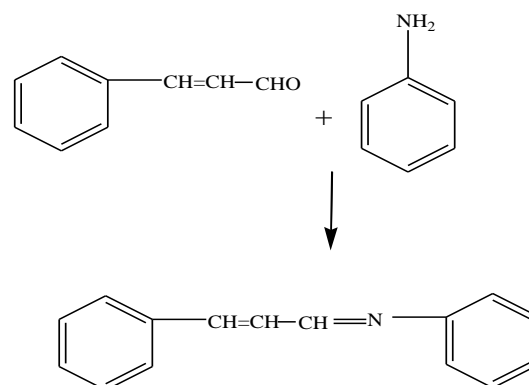
### Introduction.

The Schiff bases are also called as imines [1], anils and azomithines. The kinetic studies of Schiff base formation as well as other carbonyl addition reaction have been interested chemists for some time. Schiff bases derived from anilines and its derivative with aromatic aldehyde have a wide variety of applications in biological [2,3] and analytical chemistry[4]. Schiff bases are known to be neoplasm inhibitors [5,6], antiviral[7], anticonvulsants[8], antimicrobial [9], anticancer[10], plant growth regulator[11] and antitubercular[12], agents.

The study of kinetics of formation and hydrolysis of Schiff bases has received a considerable attention due to its relevance to the transformation (conversion) of  $>C=O$  to  $>C=N$  and vice versa in biochemical processes [13-17]. Schiff bases formation involves a two step reaction between the carbonyl compound and the amino compound. First, addition takes place to form a carbinolamine which then undergoes dehydration [18]. Both step are reversible and subject to general acid – base catalyst [19]. Literature survey reveals that a great deal of work has been reported on the complexation of metal ions with Schiff bases for the study of structure and stability of the complexes. The catalytic effect of hydrogen, hydroxyl and metal ions on the formation and hydrolysis of imines have been studied by several workers[16,17-20]1. In the present work we report here kinetic study and mechanism of the formation of Schiff base, N- Cinnemaldine aniline in ethanol medium spectrophotometrically.

### Experimental Section

All chemicals used were of AR grade. Schiff base were prepared by refluxing equimolar quantity of cinnemaldehyde and aniline in ethanol medium for about three hours. The mixture was cooled and filtered to obtain solid Schiff base and recrystallised from ethanol. The purity was checked by melting point ( $51^{\circ}C$ ) and TLC.



### Results and Discussion

Kinetics of formation of Schiff bases was carried out at equal concentration. The second order rate constant was calculated by using equation of second order at equal concentration and graphical k values determined from the straight line plots of  $1/(a-x)$  verses time.

Reaction order Effect of change in concentration of cinnemaldehyde on the rate of reaction. The kinetic study is carried out at different concentration of cinnemaldehyde ( $1 \times 10^{-3}$  to  $5 \times 10^{-3}$  mol dm<sup>-3</sup>) at constant concentration of aniline ( $5 \times 10^{-3}$  mol dm<sup>-3</sup>) in ethanol medium at 303 k. The plot of  $dc/dt$  against  $\log [A.A]$  is liner and slope of plot was found to be nearly one, indicates the order with respect to cinnemaldehyde is first order (Table1).

**Table 1. Effect of change in concentration of cinnemaldehyde on the rate of reaction.**

[C.A.] $\times 10^{-3}$ (mol dm <sup>-3</sup> )	[ANI] $\times 10^{-3}$ (mol dm <sup>-3</sup> )	k $\times 10^{-3}$ (mol <sup>-1</sup> dm <sup>3</sup> s <sup>-1</sup> )
1	5	11.04
2	5	11.18
3	5	11.28
4	5	11.48
5	5	11.58

### Effect of change in concentration of aniline on the rate of reaction.

The kinetic study is carried out at different concentration of aniline ( $1 \times 10^{-3}$  to  $5 \times 10^{-3}$  mol dm<sup>-3</sup>) at constant concentration of cinnamaldehyde ( $5 \times 10^{-3}$  mol dm<sup>-3</sup>) in ethanol medium at 303 k. The plot of dc/dt against log [ANI] is linear and slope of plot was found to be nearly one, indicates the order with respect to aniline is first order (Table2).

**Table 2. Effect of change in concentration of aniline on the rate of reaction.**

ANI]×10 <sup>-3</sup> (mol dm <sup>-3</sup> )	[C.A] [×10 <sup>-3</sup> (mol dm <sup>-3</sup> )	k ×10 <sup>-3</sup> (mol <sup>-1</sup> dm <sup>3</sup> s <sup>-1</sup> )
1	5	10.88
2	5	11.08
3	5	11.13
4	5	11.23
5	5	11.3

### Effect of temperature on the rate of reaction

Kinetic measurements were carried out at four different temperatures 303,308,313 and 318 k at constant concentration of [C.A.] ( $5 \times 10^{-3}$  mol dm<sup>-3</sup>) and aniline ( $5 \times 10^{-3}$  mol dm<sup>-3</sup>) in ethanol. The Second order rate constant depends on the reaction temperature [21,22] (Table 3).

The thermodynamic parameters like energy of activation (Ea), enthalpy of activation ( $\Delta H^*$ ), entropy of activation ( $\Delta S^*$ ), free energy ( $\Delta G^*$ ) and frequency factor (A) were calculated Table 4). From the value of thermodynamic parameter it is observed that  $\Delta H^*$  and  $\Delta S^*$  are the important parameter in controlling the rate of reactions. The negative value of entropy of activation indicates that activated complex is less probable and rate is slower. The negative values of entropy of activation show that the intermediate transition state is rigid. The relatively small values of  $\Delta H^*$  and the negative  $\Delta S^*$  values are consistent with the reactions which generally proceeds through highly organized transitions states [23]. If both the reactants are likely charged, the charge density on the surface in the transition state will be more and hence there can be increase in solvation leading again to a negative  $\Delta S^*$  [24].

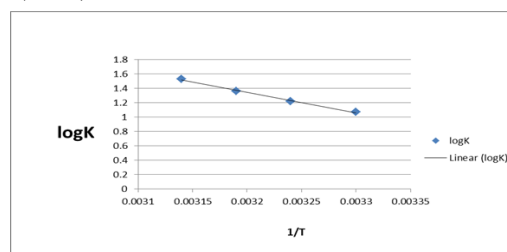
**Table 3. Effect of temperature on the rate of reaction.**

Temp.	54.571 kJ (dm <sup>3</sup> mol <sup>-1</sup> s <sup>-1</sup> )
303	11.82
308	16.5
313	23.34
318	34.45

The rate constant values of Schiff base formation at different temperature are listed in the Table-3. The rate of formation of Schiff bases increases with increasing temperature.

The plot of log k versus 1/T is straight line (Fig 1).

Schiff bases formation involves a two-step reaction between the carbonyl compound and the amino compound. First addition takes place to form a carbinolamine which then undergoes dehydration. Both steps exhibit general acid base catalysis [25, 26]. It was found that the reaction was first order in the carbonyl compounds and first order in amine



**Fig 1. The plot of log k versus 1/T**

Spectroscopic studies have revealed that, under mild acidic condition there is a fast disappearance of the carbonyl function followed by a slow appearance of the product [27]. Under mild acidic condition, rapid addition of amine to the carbonyl compound followed by the acid-catalysed dehydration of the adduct is the rate-controlling step. An increase in acidity would promote the reaction because it increases the rate of dehydration step. But with a further increase in acidity the nucleophile (RNH<sub>2</sub>) may get protonated (RNH<sub>3</sub><sup>+</sup>) and, thus deprived of its lone pair of electrons and is no longer in a position to attack the positive carbonyl carbon (Scheme.1).

**Table 4. Thermodynamic parameters**

Ea	48.571 kJ
A	$1.26 \times 10^2$
$\Delta H$	42.051 kJ
$\Delta S^*$	$-148.964 \text{ JK}^{-1} \text{ mol}^{-1}$
$\Delta G^*$	$77.946 \text{ kJ mol}^{-1}$

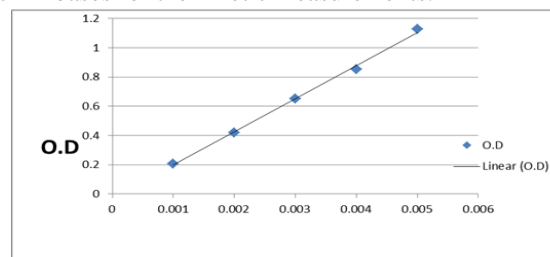
### Kinetic measurements

#### Lambert-Beer's Law

The rate of formation of Schiff base was followed by spectrophotometer at  $\lambda_{\text{max}}$ . By using UV VIS, 1601 Shimadzu spectrophotometer. The basis of the spectrophotometry is Beer's law. This relates the absorbance of a solution to the concentration of the species present. The prerequisite of the spectrophotometry is the validity of the Beer's law. It is possible to ascertain the concentration of a given species in solution if it absorbs radiation of a particular wavelength and obeys Beer's law. True variation of this law can arise when moderately concentrated solutions are used. The law is obeyed only in dilute solutions.

#### Standard curves

The solution of Schiff base of various concentrations (0.001 M to 0.005 M) was prepared using ethanol solvent. Using 0.005 M solution,  $\lambda_{\text{max}}$  was determined. The absorbance of each coloured solution was then measured, at this  $\lambda_{\text{max}}$  435 nm. The readings are recorded in Table-5. The plot of absorbance (optical density) versus concentration of the Schiff bases has been obtained as a straight line (Fig-3). The plot was used as standard curve for the determination of concentration of Schiff bases for the kinetic measurements.



**Fig 2. Standard Curve O.D Vs Concentration.**

### Experimental procedure for kinetic measurements

Cinnamaldehyde solution (20 ml) of desired molarity was taken into 50 ml flask. In another 50 ml flask amine solution (20 ml) of desired molarity were taken. Both the flasks were then allowed to stand in thermostatic water bath to attain the required temperature. Then content of the flask having amine solution transferred to the flask containing cinnamaldehyde. Thus obtained reaction mixture was thoroughly shaken and kept in thermostatic water bath at desired temperature. After mixing, the reaction mixture was transferred to a quartz cell and the increase of absorbance due to Schiff base formation with time was followed against the blank kept in another quartz cell at  $\lambda_{max}$ . At different time intervals solution was employed to determine optical density. From this optical density, concentration (x) of Schiff base present at particular time was determined with the help of the standard curve. It was observed that this method gave reproducible and quantitative result.

**Table. 5 The absorbance of each coloured solution was recorded at this  $\lambda_{max}$  420 nm.**

Conc. of S.B.	O.D.
0.001	0.208
0.002	0.419
0.003	0.651
0.004	0.852
0.005	1.126

### Methods of calculations of the rate constant

For the equimolar quantities of reactants, value of second order rate constant were calculated using relation.

$$K = \frac{1}{t} \frac{x}{a(a-x)}$$

If concentrations of reactants are different, the values of second order rate constants have been calculated by using relation

$$K = \frac{2.303 \log b(a-x)}{(a-b)t \log a(b-x)}$$

Where,

t = time in second.

a = initial concentration of cinnamaldehyde.

b = initial concentration of amine.

x = amount of Schiff base formed in time t.

Graphical k values were obtained from the slope of the linear plot of  $1/(a-x)$  versus time (t) for equal concentration and plot of  $\log [(a-x)/(b-x)]$  versus time (t) for unequal concentrations.

### Order of reaction with respect to cinnamaldehyde

To determine the order of reaction with respect to cinnamaldehyde, the reaction has been carried out at different concentrations of cinnamaldehyde by keeping the concentration of amine constant at a particular temperature. van't Hoff's differential method [28] was applied to determine the order with respect to cinnamaldehyde by equation [29,30].

$$\log dc_1/dt - \log dc_2/dt$$

$$n = \log C_1 - \log C_2$$

The amount of product formed (x) was plotted against the time in minutes. From the curve, the values of dc/dt have been calculated.

### Order of reaction with respect to amine

The order of reaction with respect to amine was determined by varying concentration of amine by keeping the concentration

of cinnamaldehyde constant at a particular temperature. The order with respect to amine was determined by applying Van't Hoff's differential method. The values of dc/dt were evaluated by plotting amount of product formed (x) against time (t) in minutes.

From the effect of temperature on the reaction rate the energy of activation  $E_a$ , enthalpy of activation ( $\Delta H^*$ ), entropy of activation ( $\Delta S^*$ ), free energy ( $\Delta G^*$ ) and frequency factor (A) were calculated.

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